Kinetics and Thermodynamic Studies of Adsorption of Methylene Blue from Aqueous Solutions onto Paliurus spina-christi Mill. Frutis and Seeds

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Abstract: In the present study, Paliurus spina-christi Mill. Frutis and Seeds (PSCFS) was used as an adsorbent for the investigation of the adsorption kinetics and thermodynamic parameters of methylene blue (MB) from aqueous solution at various temperatures and concentrations. Six kinetic models, the pseudo-first-order, pseudo-second-order, intra-particle diffusion, Elovich, Bangham and Avrami equations were used to predict the adsorption rate constants. It was found that the kinetics of the adsorption of MB onto PSCFS at different operating conditions was the best described by the pseudo-second-order model. The rate parameters of the other kinetic models for adsorption were also evaluated and compared to identify the adsorption mechanisms. The equilibrium constants were used to calculate thermodynamic parameters, such as the change of free energy, enthalpy and entropy. Thermodynamic analysis showed that adsorption was favourable and spontaneous, endothermic physical adsorption and increased disorder and randomness at the solid-solution interface of MB onto PSCFS.

Keywords: Adsorption, Thermodynamic, Kinetic, Plant, Raw material

I. Introduction

Dye removal from wastewater is a topic of great interest in water pollutions which is a serious environmental damage. Methylene blue is the most soluble dye in water in general used for printing cotton and tannin, dyeig leather, antiseptic and medical purposes [1]. Cationic dyes like methylene blue (MB) are more toxic than anionic dyes [2]. Although this dye does not harm people very much, it can cause irritation to the eyes and the skin [3]. MB may cause vomiting, nausea, diarrhea, profuse sweating, mental confusion, and gastritis and it can make breathing difficult [4]. Therefore, it has been focused on the removal of this dye from the wastewater. Numerous methods such as reverse osmosis, ion exchange, membrane filtration, conventional coagulation, chemical precipitation, and adsorption have been generally used to remove dyes from wastewater [5,6]. Of the removal methods, adsorption is one that sees increasing interest due to its potential efficiency, low energy consumption, high selectivity at molecular level, easy operation, and ability to separate various chemical compounds [7, 8]. Adsorption is a physical-chemical treatment of wastewater in which the dissolved molecule is attached to an adsorbent surface by means of physical and chemical properties. Depending on the nature of the adsorbent and the origin of the dyes, different interactions may be performed, such as electrostatic interaction and Van der Waals forces [9]. One of the important parameters in the adsorption process is the cost of the adsorbent. For this reason, it is important to use low cost adsorbents. Industrial by-products and agricultural waste products form important low-cost sources of adsorbent material. In recent years, a large number of low-cost adsorbents have been used to remove water from waste water such as cucumber peels [10], meranti sawdust[11], bagasse [12], durian leaf powder[13], watermelon seed hulls [10], grape pulp [14], chitosan [15], kenaf core fibers [16], hazelnut Shell [17], Delonix regia plant leaf [18], etc. However, no adsorption process has been reported on the use of PSCFS as adsorbent for the removal of cationic dye (MB) from water solution. In this study, an inexpensive natural adsorbent, PSCFS, was used to remove from the aqueous solution of MB. Contact time, initial concentration, initial pH and temperature effect on MB adsorption were investigated. In this article, kinetic and thermodynamic parameters were determined using available parameters.

II. Materials and methods

2.1 Adsorbent

PSCFS were collected from of Manyas region Balikesir city, Turkey. The samples were washed several times with distilled water to remove impurity such as sand and dirt, and then dried in an oven at 55 °C for 24 h. Dried samples were cut into pieces and then ground by using a crusher. The powder were passed through sieves and particles were collected under 0.06 mm. These samples were used as adsorbent for adsorption experiments.
2.2 Adsorbate and other chemicals

The basic cationic dye, MB, was used as the adsorbate in this study. The formula of methylene blue is C_{6}H_{18}N_{3}SCl.3H_{2}O, and molar weight 319.85 g. The chemical structure of this dye is shown in Fig. 1. This dye was taken from Merck and used without any purification. The stock dye solution was prepared by dissolving 1 g of methylene blue in 1000 ml distilled water. The solutions used for adsorption process were obtained by diluting the stock dye solution with deionized water to give the appropriate concentration of the solutions.

The T80+ UV/VIS spectrophotometer was used to determine the concentration of MB in solution at a maximum absorption wavelength of 660 nm. A calibration curve was plotted between absorbance and certain concentrations of dye solution. Unknown MB concentration was measured using calibration curve.

![Chemical structure of methylene blue](image)

Fig. 1. Chemical structure of methylene blue

2.3 Adsorption experiment

Adsorption measurement was determined by batch experiments of known amount of the adsorbent with 1000 ml of aqueous methylene blue solutions of known concentration in a series of 250 ml conical flasks. The effect of the contact time on the amount dye adsorbed were investigated at different initial concentration of dye (20, 30, 40 and 50 mg/L) at different temperatures (298 K, 308 K and 318 K). 4 g PSCFS were added to 1 liter of dye solutions prepared at certain concentrations, natural pH and desired temperatures at 150 rpm for specific time intervals. The concentration of dye remaining in solution without adsorption was determined by UV spectrophotometer using calibration curve. Adsorption experiments were made considering natural solution pH, contact time, initial methylene blue dye concentration and temperature for adsorption kinetic and thermodynamic study. The amount of dye adsorbed onto PSCFS at time \( t \) is \( q_t \) (milligrams per gram) which was calculated by the following mass balance equation:

\[
q_t = \frac{(C_o - C_t)}{m} V
\]

where \( C_o \) is the initial dye concentration (milligrams per liter), \( C_t \) is the concentration of dye at any time \( t \), \( V \) is the volume of solution (liters), and \( m \) is the mass of adsorbent in grams.

III. Theory

Kinetic models are used to check experimental data from the adsorption of MB onto PSCFS. The kinetics of dyes adsorption onto PSCFS are important to select the best test conditions for the adsorption process with the batch technique. The kinetic parameters, which are useful for the prediction of adsorption rate, give important information for designing and modeling the adsorption processes. In this study, the kinetics of MB onto PSCFS was analyzed using pseudo-first order, pseudo-second order, intra-particle diffusion, Elovich, Bangham and Avrami kinetic models.

3.1 Pseudo-first order model

Lagerrn’s kinetics equation may have been the first one in describing the adsorption of liquid-solid systems based on solid capacity. In order to distinguish kinetics equation based on concentration of solution and adsorption capacity of solid, Lagerrn’s first order rate equation has been called pseudo-first order[ 19]. The linear form of the pseudo-first order model is given as:

\[
\log(q_m - q_t) = \log q_m - \frac{k_1 t}{2.303}
\]

where \( q_m \) and \( q_t \) (mg/g) are the adsorption capacities at equilibrium and at time \( t \) respectively. \( k_1 \) (min\(^{-1}\)) is the rate constant of pseudo-first order adsorption. In order to obtain the constants of this model, the straight line plots of \( \log(q_m - q_t) \) against \( t \) are drawn. The constants determined from the slope and intercept of the plot.
Kinetics And Thermodynamic Studies Of Adsorption Of Methylene Blue From Aqueous Solution...

3.2 Pseudo-second order model

The linear form of this model based on adsorption capacity is given as [19]:

\[
\frac{t}{q_t} = \frac{1}{k_2q_m^2} + \frac{1}{q_m} t
\]

(3)

and

\[
h = k_2q_m^2
\]

(4)

Where \( k_2 \) is the rate of sorption (g/mg min), \( q_m \) the amount of adsorbate adsorbed onto adsorbent at equilibrium (mg/g), \( q_t \) is the amount of dye adsorbed at any time (mg/g) and \( h \) is the initial sorption rate (mg/g min). The plot of \( t/q_t \) versus \( t \) shows a linear relationship. Values of \( k_2 \) and equilibrium adsorption capacity \( q_e \) were calculated from the intercept and slope of the plot of \( t/q_t \) versus \( t \) according to Eq. 2.

3.3 Intraparticle diffusion model

In intraparticle diffusion model, the adsorbate transportation from the solution phase to the surface of the adsorbent particles occurs in several steps. All adsorption processes can be controlled by one or more steps, e.g., film or external diffusion, pore diffusion, surface diffusion and adsorption on the pore surface, or a combination of more than one step. The linear form of the intraparticle equation as follow:

\[
q_t = k_i t^{1/2} + C
\]

(5)

where \( C \) is the intercept and \( k_i \) is the intra-particle diffusion rate constant (mg/g min\(^{1/2}\)). When intraparticle diffusion alone is the rate limiting step, then the plot of \( q_t \) versus \( t^{1/2} \) passes through the origin [20]. When film diffusion is also taking place then the intercept is \( C \), which gives the idea on the thickness of the boundary layer.

3.4 Elovich kinetic model

Elovich equation is also used successfully to describe second order kinetic assuming that the actual solid surfaces are energetically heterogeneous, but the equation does not propose any definite mechanism for adsorbate-adsorbent [21]. It has extensively been accepted that the chemisorption process can be described by this semi-empirical equation [22]. The linear form of this equation is given by:

\[
q_t = \frac{\ln(a_f b_E)}{b_E} + \frac{1}{b_E} \ln t
\]

(6)

Where \( a_f \) is the initial adsorption rate (mg/g min), and the parameter \( b_E \) is related to the extent of surface coverage and activation energy for chemisorption (g/mg). The Elovich coefficients could be computed from the plots \( q_t \) versus \( \ln t \). The initial adsorption rate \( (a_f) \) and the desorption constant \( (b_E) \) were calculated from the intercept and slope of the straight-line plots of \( q_t \) against \( \ln t \).

3.5 Bangham kinetic model

Bangham’s equation was used to evaluate whether the adsorption is pore-diffusion controlled. Bangham’s plot should be linear if intraparticle diffusion is the only rate controlling step [23]. Non-linearity of the Bangham plots as well as the significant intercept values obtained from the intraparticle diffusion model indicated that both film diffusion as well as pore diffusion were rate-limiting [24].

\[
\log \log \left( \frac{C_o}{C_o - q_t} \right) = \log \left( \frac{k_B M}{2.303V} \right) + b \log t
\]

(7)

Where \( C_o \) is initial concentration (mg/L), \( V \) is volume of the solution (mL), \( M \) is weight of the adsorbent (g/L), \( q_t \) is amount of adsorbate retained at time \( t \) (mg/g) and \( b \), \( k_B \) are constants.

3.6 Avrami kinetic model

The Avrami kinetic equation determines some kinetic parameters, as possible changes of the adsorption rates in function of the initial concentration and the adsorption time, as well as the determination of fractionary kinetic orders [25]. The linearized form of Avrami equation is shown as follows:

\[
\ln \left( \ln \left( \frac{q_m}{q_m - q_t} \right) \right) = n_{Av} \ln k_{Av} + n_{Av} \ln t
\]

(8)

Where \( k_{Av} \) is the Avrami adsorption kinetic constant and \( n_{Av} \) is another constant, which is related to the adsorption mechanism changes. The slopes and intersections values of this equation provide the \( n_{Av} \) and \( \ln k_{Av} \) values, respectively. \( n_{Av} \) value can be used to verify possible interactions of the adsorption mechanisms in relation to the contact time and the temperature.
IV. Thermodynamic parameters

Thermodynamic investigation is required to determine whether the adsorption process is favorable. The thermodynamic parameters, namely free energy (ΔG°), enthalpy (ΔH°), and entropy (ΔS°) have an important function to determine heat change the adsorption process for dye and PSCFS. These parameters are calculated by the following equations:

\[ K_e = \frac{C_{Ads}}{C_e} \quad (9) \]

\[ \ln K_e = \frac{\Delta S^o}{R} - \frac{\Delta H^o}{RT} \quad (10) \]

\[ \Delta G^o = \Delta H^o - T\Delta S^o \quad (11) \]

where, \( K_e \) is the equilibrium constant, \( C_{Ads} \) is the amount of dye adsorbed mg on the PSCFS per dm³ of the solution at equilibrium, adsorbent adsorbent per unit liter of solution and \( C_e \) is the equilibrium concentration of dye in the solution (milligrams per liter). \( R \) is the universal gas constant (8.314 J/mol K) and \( T \) is the temperature (Kelvin). \( \Delta H^o \) and \( \Delta S^o \) parameters are calculated from the slope and intercept of the plot \( \ln K_e \) vs. 1/\( T \).

V. Results and discussion

5.1 Effect of contact time on MB

Contact time is one of the physical parameters used economically for the design of wastewater treatment plants [26]. Fig. 1 shows that the removal of dye from the solutions is rapid in the beginning period and that the velocity near the reach of the balance decreases. At the beginning, the surface of the adsorption process is large, so the adsorption to this surface is fast. The equilibrium time in dye adsorption was determined to be 140 min.

![Fig. 1. Effect of contact time and initial dye concentration on dye removal onto PSCFS (natural initial pH, at 318 K, agitation speed 150 rpm).](image)

5.2 Effect of initial dye concentrations

Figure 2 shows that the increase in the initial dye concentration caused the increase in adsorption capacity. As the initial dye concentration increases from 20 to 50 mg/L, the adsorption capacity of dye onto PSCFS from 16.43 to 37.80 mg/g. These data indicate that the initial dye concentration plays an important role in the dye adsorption capacity. However, as shown in Figure 2, the percent removal of dye decreased from 82.15 to 75.59 % on increasing the initial dye concentration from 20 to 50 mg/L. As seen from the figure, the amount of adsorbed dye increases and the percentage of adsorption decreases with the initial dye concentration increases. This indicates that the initial dye concentration provides a driving force to overcome the resistance to dye transfer between the solid phase and the water phase. The increase in initial concentration increases the interaction between adsorbent and dye. Therefore, an increase in initial dye concentration leads to increase in
the adsorption amount of dye. As can be seen from the experimental data obtained, the adsorption of MB on PSCFS contain three stages; initially fast adsorption, followed by slower adsorption and eventually no adsorption. Similar types of results are reported by various researchers for methylene blue adsorption on sawdust, rice husk, spent tea leaves [27], on seed watermelon pulp [28], on corn husk [29], on sunflower seed hull [30], on raw olive pomace [31] and on maize silk powder [32].

5.3 Effect of temperature on MB dye adsorption

The effect of temperature is an important physical parameter since it changes the adsorption capacity of the adsorbent. To study the effect of temperature on the adsorption of MB dye adsorption onto PSCFS, the experiments were carried out at temperatures of 25, 35 and 45 °C. Figure 3 shows the effect of temperature on the adsorption of by MB onto PSCFS. As you can see, the amount of adsorbed dye and the percentage of adsorption onto PSCFS were found to increase with increasing temperature. This shows that surface activity increases at higher temperatures and adsorption is endothermic.

5.6 Thermodynamic parameters

Thermodynamic parameters have been investigated to determine whether the adsorption phenomenon is favorable or not. ΔG°, ΔH°, and ΔS° were obtained from the experimental studies at different temperatures of 298, 308 and 318 K by the applying of Eqs. 8, 9 and 10 and plot which is not shown here. All thermodynamic parameters are tabulated in Table 1. The ΔG° values change from -2.4334 to -3.7478 while the temperature varies from 298 to 318. These negative values indicate that dye adsorption reaction is spontaneous in nature at all the studied temperatures. The increase in ΔG° values with increasing temperature is an indication that MB adsorption on PSCFS is more favorable. The positive ΔH° value indicates that the reaction is endothermic and consumes energy in the adsorption process. The positive value of ΔS° indicated increased randomness at the solid-solute interface during adsorption.

![Fig. 2. Effect of initial MB dye concentration on the adsorption of MB onto PSCFS (volume of dye solution 1 L, temperature 45 °C, PSCFS dosage 4 g, natural initial pH, agitation speed 150 rpm).](image)

![Fig. 3. Effect of temperature on the adsorption of MB onto PSCFS (volume of dye solution 1 L, initial dye concentration 50 mg/L, PSCFS dosage 4 g, natural initial pH, agitation speed 150 rpm).](image)
### Table 1 Thermodynamic parameters for adsorption of MB onto PSCFS at different temperatures and concentrations

<table>
<thead>
<tr>
<th>Temperature (K)</th>
<th>$-\Delta G^\circ$ (kJ/mol)</th>
<th>$\Delta S^\circ$ (J/mol K)</th>
<th>$\Delta H^\circ$ (kJ/mol)</th>
</tr>
</thead>
<tbody>
<tr>
<td>298</td>
<td>2.4334</td>
<td>66.9117</td>
<td>17.5301</td>
</tr>
<tr>
<td>308</td>
<td>3.0787</td>
<td></td>
<td></td>
</tr>
<tr>
<td>318</td>
<td>3.3478</td>
<td></td>
<td></td>
</tr>
</tbody>
</table>

5.7 Adsorption kinetics

The study of kinetics in adsorption treatment from water solutions is important because it gives important information about the reaction and the mechanism of the adsorption process. In this study, the pseudo-first-order, pseudo-second-order, intra-particle diffusion, Elovich, Bangham and Avrami kinetics model was tested for the adsorption of MB into PSCFS. The best fit model was chosen based on linear regression coefficient of correlation coefficients $R^2$ values. These models have been investigated according to experimental data at different temperatures and concentrations. The linear graphs for these models are shown for only 298 K temperature. All kinetic parameters obtained from fitting model plots with experimental data under various conditions are presented in Tables 2, 3 and 4.

**Fig. 4.** Pseudo-first-order kinetic model on different initial dye concentration at 298 K.

The $R^2$ values obtained for the pseudo-first-order kinetic model were lower than that of the pseudo-second-order kinetic model and the experimental $q_e$ values did not agree with the calculated values obtained from the linear plots (Table 2, 3 and 4). This indicates that the adsorption of MB onto PSCFS does not follow pseudo-first-order kinetics. The similar results were found for the adsorption of MB ions on various adsorbents by several authors [33, 34].

**Fig. 5.** Pseudo-second-order kinetic model on different initial dye concentration at 298 K.
When an evaluation is made taking into consideration the temperatures and concentrations used for the kinetic analyzes, it is found that the maximum average $k_2$ value is $6.9475 \times 10^{-3}$ g / mg min at 318 K. It was determined that the maximum initial velocity value was 0.6623 mg / g min at a concentration of 318K 50 mg / L. When theoretically and experimentally calculated adsorption capacities are compared, it is determined that the values obtained at 318 K are closer to each other. When the correlation coefficients are examined, it is observed that the most suitable value is in this model. This shows that it is more appropriate to explain the obtained data with a pseudo-second-order kinetic model.

![Intraparticle diffusion kinetic model on different initial dye concentration at 298 K.](image)

**Fig.6.** Intraparticle diffusion kinetic model on different initial dye concentration at 298 K.

When Figure 6 is examined, it is observed that there are three phases in MB adsorption. It can be interpreted that the dye molecules are sorption to the outer surface of the adsorbent, in the first stage when the amount of adsorbed material increases rapidly, diffuse in the pores of the adsorbent in the second stage where the adsorption rate slows down and in the third stage the adsorption phenomenon is balanced. When the data related to intraparticle diffusion in Tables 2, 3 and 4 were examined, it was observed that $R^2$ values were small. From tables data, with increasing temperature and concentration, it is observed that C values increased and $K_p$ values decreased. The linear plots do not pass through the origin which indicates that intraparticle diffusion is not the only rate controlling step [35].

![Elovich kinetic model on different initial dye concentration at 298 K.](image)

**Fig.7.** Elovich kinetic model on different initial dye concentration at 298 K.

The Elovich equation is a very common model for describing chemisorption processes. The constants of Elovich equation and the correlation coefficient are given in tables 2, 3 and 4. It is observed that the values of $a_0$ (mg g$^{-1}$ min$^{-1}$) increase with the increase of initial concentration and temperature in the range of studied
values. The values of the regression coefficient ($R^2 = 0.8462-0.9540$) of the Elovich kinetic model suggest that kinetic data did not follow the Elovich model.

Tables 2, 3 and 4 are given the kinetic constants obtained from the Bangham equation. When the tables are examined, it is seen that $b$ value decreases with increasing temperature. It is observed that $k_0$ value increases with increasing temperature. It is seen that the correlation coefficient obtained from the tables Elovich model is larger than that obtained by Bangham model. These results confirm that the adsorption rate is not controlled only by the pore diffusion. The graphs of the kinetic model of Avrami at various concentrations and 298 K are shown in fig.9. The kinetic data calculated from the Avrami equation are given in Tables 2, 3 and 4. From the tables, it seems that the calculated Avrami constants depend on the temperature. From these constants, while the values of $n_{Av}$ decreases with temperature, the values of $k_{Av}$ increase with temperature. When the temperature rises, the increase in $k_{Av}$ value is indicative of increased the reaction speed with temperature.
Table 2 Kinetic models for the adsorption of MB adsorption onto PSCFS at 298 K

<table>
<thead>
<tr>
<th>Kinetic models</th>
<th>Parameters</th>
<th>Adsorbent matter</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>qₑ(0) (mg/g)</td>
<td>20 mg/L</td>
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<td>Pseudo-first order</td>
<td>qₑ(0) (mg/g)</td>
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<td></td>
<td>k₁ (min⁻¹)</td>
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<tr>
<td></td>
<td>R²</td>
<td>0.9065</td>
</tr>
<tr>
<td>Pseudo-second order</td>
<td>qₑ(0) (mg/g)</td>
<td>5.7625</td>
</tr>
<tr>
<td></td>
<td>qₑ(0) (mg/g)</td>
<td>5.1256</td>
</tr>
<tr>
<td></td>
<td>k₂ (g/mg min)</td>
<td>3.43x10⁻³</td>
</tr>
<tr>
<td></td>
<td>h (mg/g min)</td>
<td>0.0901</td>
</tr>
<tr>
<td></td>
<td>R²</td>
<td>0.8604</td>
</tr>
<tr>
<td>Intraparticle diffusion</td>
<td>kₐ (mg/g min)</td>
<td>0.2713</td>
</tr>
<tr>
<td></td>
<td>C (mg/g)</td>
<td>0.5121</td>
</tr>
<tr>
<td></td>
<td>R²</td>
<td>0.7478</td>
</tr>
<tr>
<td>Avrami</td>
<td>nₐₐ</td>
<td>1.4964</td>
</tr>
<tr>
<td></td>
<td>kₐ (mg/g min)</td>
<td>2.32x10⁻²</td>
</tr>
<tr>
<td></td>
<td>R²</td>
<td>0.9878</td>
</tr>
</tbody>
</table>

Table 3 Kinetic models for the adsorption of MB adsorption onto PSCFS at 308 K

<table>
<thead>
<tr>
<th>Kinetic models</th>
<th>Parameters</th>
<th>Adsorbent matter</th>
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<tr>
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<td>qₑ(0) (mg/g)</td>
<td>20 mg/L</td>
</tr>
<tr>
<td>Pseudo-first order</td>
<td>qₑ(0) (mg/g)</td>
<td>5.3432</td>
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<td></td>
<td>k₁ (min⁻¹)</td>
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<td></td>
<td>R²</td>
<td>0.8849</td>
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<td>Pseudo-second order</td>
<td>qₑ(0) (mg/g)</td>
<td>3.9975</td>
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<td>qₑ(0) (mg/g)</td>
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<td>k₂ (g/mg min)</td>
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<td>h (mg/g min)</td>
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<td>R²</td>
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<td>Intraparticle diffusion</td>
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<td></td>
<td>C (mg/g)</td>
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<td></td>
<td>R²</td>
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<td>Avrami</td>
<td>nₐₐ</td>
<td>1.90x10⁻²</td>
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<td></td>
<td>kₐ (mg/g min)</td>
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<tr>
<td></td>
<td>R²</td>
<td>0.9878</td>
</tr>
</tbody>
</table>

When comparing R² values obtained from kinetic models, the results reveal that kinetic models fitted in the order: pseudo-second order>Avrami>pseudo-first order>Elovich>Bangham>intraparticle diffusion. This demonstrates that the pseudo-second-order kinetic model better describes the adsorption process. Moreover, the qₑ values obtained with this model are closer to the values obtained from the experimental data.

DOI: 10.9790/5736-1005015363  www.iosrjournals.org  61 [Page]
Implied that the reaction is spontaneous in nature, which is also confirmed by the positive values of enthalpy change. The maximum removal of dye at 318 K was found to be 82.

The adsorption of dye MB on PSCSF was examined at different experimental conditions. The results show that adsorption increases with increase in temperature, initial concentration and contact time.

VI. Conclusions

1. The adsorption of dye MB on PSCSF was examined at different experimental conditions. The results show that adsorption increases with increase in temperature, initial concentration and contact time.

2. The maximum removal of dye at 318 K was found to be 82.

3. Kinetic study shows the adsorption reaction follows pseudo-second order kinetic model (average R² value is 0.9697).

4. The negative value of change in ΔG° (Gibbs’s free energy) implied that the reaction is spontaneous in nature, the values are more negative with temperature that adsorption is favored with increased in temperature for the studied adsorbent.

5. Adsorption is favored at high temperature fort he studied adsorbent which is also confirmed by the positive values of enthalpy change.

Table 4 Kinetic models for the adsorption of MB adsorption onto PSCSF at 318 K

<table>
<thead>
<tr>
<th>Kinetic models</th>
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<th>20 mg/L</th>
<th>30 mg/L</th>
<th>40 mg/L</th>
<th>50 mg/L</th>
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<td>Pseudo-first order</td>
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<td>8.7640</td>
<td>8.6000</td>
<td>12.3197</td>
</tr>
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<td></td>
<td>k1 (min⁻¹)</td>
<td>3.80x10⁻¹</td>
<td>3.52x10⁻²</td>
<td>3.66x10⁻²</td>
<td>4.26x10⁻²</td>
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<tr>
<td></td>
<td>R²</td>
<td>0.8798</td>
<td>0.8502</td>
<td>0.9684</td>
<td>0.9350</td>
</tr>
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<td>Pseudo-second order</td>
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<td>6.0295</td>
<td>7.6125</td>
<td>9.4475</td>
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<td>qe,cal (mg/g)</td>
<td>4.6019</td>
<td>6.8776</td>
<td>8.3403</td>
<td>10.2987</td>
</tr>
<tr>
<td></td>
<td>k2 (g/mg min)</td>
<td>8.96x10⁻³</td>
<td>5.53x10⁻³</td>
<td>7.06x10⁻³</td>
<td>6.24x10⁻³</td>
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<td>h(mg/g min)</td>
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<td>R²</td>
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<td>Intraparticle diffusion</td>
<td>k (mg/g min⁻¹)</td>
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<td>C (mg/g)</td>
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<td>Avrami</td>
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References


DOI: 10.9790/5736-1005015363 www.iosrjournals.org 62 | Page
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