

Investigation of the Anodic Dissolution of Zinc in Sodium Chloride Electrolyte – A Green Process

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Abstract: The anodic dissolution of zinc electrodes in sodium chloride aqueous solution has been investigated experimentally. The effects of application of polarity reversal (PR), ultrasonic (US) enhancement, stirring, current density (CD), concentration and pH of the supporting electrolyte, and temperature of the bath were studied. The results revealed that application of PR increased the dissolution of Zn but the current was low. However, the application of US enhancement led to higher zinc dissolution accompanied with higher current efficiency (CE). The combination of US enhancement and stirring led to more dissolution of zinc. Increasing the current density and concentration of NaCl increased the dissolution of zinc and the current efficiency was almost constant. On the other hand, pH of the bath did not play a significant effect on the amount of the dissolved zinc or current efficiency. It was also observed that increasing the temperature from 10 °C to 40 °C led to a significant increase in the mass of the dissolved zinc and CE; but the increase of temperature from 40 °C to 50 or 60 °C, however, did not have a significant effect.

Keywords: Anodic dissociation, Electrochemistry, Green processes, Zinc, Zinc Oxide.

I. Introduction

Inorganic metal compounds are very important in chemical industries especially in plating industry, painting, pigments production, polishing materials production, cosmetics industries, and other industries. Chemical dissolution of metals is considered as the most important step in the production of the pure metal salts [1]. However, the use of active reagents currently utilized in the dissolution process has several disadvantages, such as: production of vapors or fumes, which requires complicated and troublesome trapping arrangements and neutralizations, extensive energy consumption, long time for metal dissolution, and the use of special materials of construction for equipment due to the strong corrosiveness of the dissolution medium. In order to overcome these disadvantages, electrolytic metal dissolution may introduce a realistic solution.

The process of selective anodic dissolution of a metal does not have the disadvantages mentioned above. In essence, the metal is treated electrolytically as an anode, by utilizing a rapid-action electrolyte of a composition which is substantially independent of the duration of the operation of the dissolution. Several inorganic compounds may be produced by the above mentioned anodic dissolution process, such as zinc sulfate, zinc oxide, zinc chloride, copper sulfate, copper oxide, copper cyanide, nickel sulfate... etc. Each of these compounds has wide industrial applications [2].

Zinc is ranked as the 23rd most common elements in the earth's crust, amounting to 0.013%. However, it ranks fourth among the metals in worldwide production and consumption [3]. The uses of zinc can be divided into six major categories: (i) coatings; (ii) casting alloys; (iii) alloying element in brass and other alloys; (iv) wrought zinc alloys; (v) zinc oxide; and (vi) zinc chemicals [4]. The most important application of zinc is a coating for steel corrosion protection. In addition, zinc is an important component in paints, cosmetics, pharmaceuticals, storage batteries, electrical equipment and an endless list of other capital applications [5–7]. As a result of its importance, the zinc corrosion and electrochemistry has been studied extensively [8–19] and reviewed thoroughly up to year 1996 by Zhang [4]. Moreover, many investigations have been conducted to determine the kinetic and thermodynamic characteristics of zinc electrode system. Extensive bibliographies of this work appear in the literatures [10,16]

The dissolution of zinc was studied extensively through the last four decades [8,18-50]. The anodic dissolution of zinc has been investigated in different electrolyte media. Some examples are: in potassium hydroxide solutions [22–29], in potassium hydroxide in presence of sodium metasilicate [30], in potassium hydroxide in presence of carbonate ions [31], in potassium hydroxide in presence of ZnO [32], in sodium hydroxide solutions [33–35], in sodium hydroxide in presence of sodium silicate and sodium tetraborate [13], in sodium hydroxide in presence of sulfur containing ions such as Na₂SO₄, Na₂SO₃, Na₂S, Na₂S₂O₃ or NH₄SCN [36], in saturated zinc sulfate solution [37], in aerated sodium sulfate solutions [38,39], in NH₄Cl and/or NH₄Cl/ZnCl₂ solutions [40,41], in NH₄Cl + NiCl₂ and NH₄Cl + NiCl₂ + ZnCl₂ electrolytes [42], in H₃BO₃ / NH₄Cl / Na₂SO₄ [43], in aqueous-methanolic trichloroacetic acid solutions [44], in ammoniacal ammonium

chloride system [45], in potassium nitrate solution [46], in aqueous salt solutions such as Cl^- , Br^- , I^- , Ac^- , SO_4^{2-} and NO_3^- [47], in sulfamic acid-formamide solution [48], in EDTA solution [49], in sodium borate solutions with or without Na_2SO_4 , $\text{Na}_2\text{S}_2\text{O}_3$ or Na_2S as aggressive agents [50], solutions saturated in with carbon dioxide under elevated pressure [51], and even in tap water [52].

The corrosion and electrochemical studies of zinc dissolution in sodium chloride solutions has been investigated in the literatures [14,17–19]. For $\text{pH} > 12$, a passive layer composed of zinc hydroxide and/or zinc oxide is formed on the surface of the zinc sample [19]. In an aerated acidic NaCl solution, the complexity of the layer formed on zinc sample exposed to NaCl media makes the corrosion mechanism of zinc difficult to be understood. It forms a complex layer composed from zinc oxide, zinc hydroxide and zinc hydroxide chloride or simonkolleite; $\text{Zn}_5(\text{OH})_8\text{Cl}_2 \cdot 2\text{H}_2\text{O}$. The zinc hydroxide chloride and zinc oxide are the dominant corrosion products [53]. In ambient CO_2 levels, zinc hydroxide carbonate or hydrozincite; $(\text{Zn}_5(\text{OH})_6(\text{CO}_3)_2) \cdot \text{H}_2\text{O}$ were additionally observed [54], with the major compounds being zinc hydroxide chloride and zinc hydroxide carbonate [55].

On the other hand, the chemical applications of ultrasound (US), "sonochemistry", has become an exciting new field of research during the past decades [56]. Moreover, the application of US in electrolytic cell has been reported to help in depassivation of the electrodes with a huge effect on the mass transport at the electrode surface [57].

In order to design an industrial unit or even a pilot unit for the electrolytic production of zinc compounds, the operating conditions of the electrolytic anodic dissolution of the zinc metal sources of these compounds have to be thoroughly investigated. In this paper, the effects of the application of US and polarity reversal (PR) and other operating parameters such as current density, concentration of electrolyte, pH, temperature, and stirring on the anodic dissolution of zinc in NaCl are presented.

II. Experimental

2.1 Surface pretreatment

All chemicals used in surface treatment were of analytically pure grades. Six operations were carried out on the surface of zinc substrate. These operations were: mechanical polishing with two different sand papers down to 4/0 grade, degreasing with solvent, degreasing with alkali solution, pickling, and chemical polishing. These operations were done in order to obtain high smooth surface free from oil, grease and scales and to avoid defects such as peeling, blistering, or poor distribution of the metal on the surface of the electrode. To remove grease, oils, waxes and fats introduced to the metal either from fabrication, stamping, pressing, or polishing, degreasing was done by using solvent cleaners and alkali soaking cleaners. Degreasing with solvent was accomplished using acetone, which possesses all the requirements necessary to ideal degreasing practices because of its ease of application and its great penetrating power. As for alkali soaking, a cleaner solution consisting of sodium hydroxide 50 g/l, sodium carbonate 20 g/l, trisodium phosphate 20 g/l, and sulfonic acid 2 g/l dissolved in bi-distilled. Degreasing was assured by washing the metal with water where, the presence of a continuous unbroken film of water over the zinc substrate indicates the complete removal of greases. To remove oxides and scales, the zinc substrate was pickled with a solution consisting of 400 g/l nitric acid and 5 g/l hydrofluoric acid in bi-distilled water. Finally, to ensure smooth surface, the zinc substrates were chemically polished by immersing the substrates in 100g/l oxalic acid solution and were allowed to boil for 5 minutes.

2.2 Preparation of electrolyte solution

All materials used in the electrolyte preparation for the anodic dissolution of zinc metal were of analytical grade chemicals. The sodium chloride supporting electrolyte was weighed and dissolved in bi-distilled water to the desired concentration and then pH adjustment was carried out to make the solution ready for the electrolysis process. Hydrochloric acid and sodium hydroxide were used for the pH adjustment. Zinc metal for the dissolution process was of high purity of 99.99%. This zinc was imported from Germany.

2.3 Experimental setup

Fig. (1) shows the inside details of the electrolytic cell.

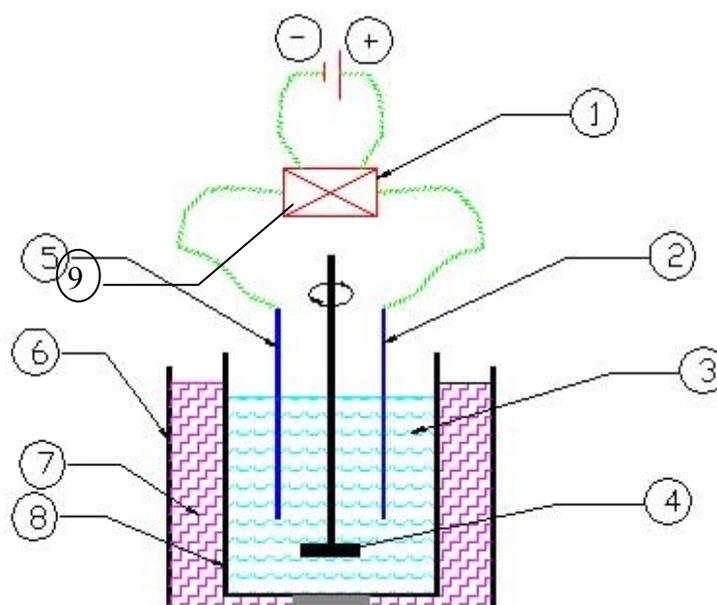


Fig. (1): Details of the experimental setup for the anodic dissolution of zinc. 1- Polarity reverser, 2- Zinc anode, 3- Supporting electrolyte, 4- Mechanical stirrer, 5- Zinc cathode, 6- Source of ultrasonic waves, 7- Water bath, 8- Glass container, 9- D.C. power supply.

The setup consisted of:

- D.C. power supply: A regulated D. C. power source (Lodestar 8203) capable to supply up to 10 A and 30 V was used to supply electric power. During each experiment the current could be kept constant by manual regulation.
- Mechanical stirrer: had a fixed speed of 200 rpm.
- Polarity reversal: a back-switch polarity reversal piece used to reverse electrode polarity during the electrolysis process. The reversal time was adjusted manually.
- pH meter: (type HI 8417 made by HANNA Instruments) of a resolution 0.01 was used to measure the pH of the electrolyte solution.

2.4 Experimental procedure

The electrolysis cell was cleaned, washed and filled with the electrolyte to the desired level to ensure complete submerge of the electrodes. In this cell, zinc substrate was used as anode. During some experiments the direct current pathway was reversed to the opposite direction and consequently the cathode was made of zinc metal. The reversal of the polarity was expected to help in the prevention of the formation of a passive film on the surface of the anode. In addition, the use of US was expected to repel off the formed passive layer to the bulk of the solution. All of the above may minimize the current drop and maintain the current efficiency of the zinc dissolution as high as possible. The zinc substrate was weighed before and after each dissolution experiment and from the weight difference the weight of the dissolved zinc metal was determined.

III. Results and Discussion

3.1 Anodic dissolution of Zn without the application of PR, US or stirring

Fig. (2) shows the change of the current passing in the cell at a potential of 2.8 V without the application of US, PR or stirring. The electrolyte consisted of 90 g/dm³ (~1.5 mole/L) sodium chloride solution. The pH was adjusted using sodium hydroxide to value of 13. The time of electrolysis was 120 minutes. The average current passing through the cell was calculated, from the area under the curve divided by 120 minutes, as 0.436 A. The dissolved Zn was 1.0634 g, and the current efficiency (CE) was 99.94%. The high CE may be attributed to the small value of the average current passing at this potential, and the dissolution was significant at the early period after the application of potential. After this effective period, the passing current decreased due to the formation of the passive layer on the surface of the anode.

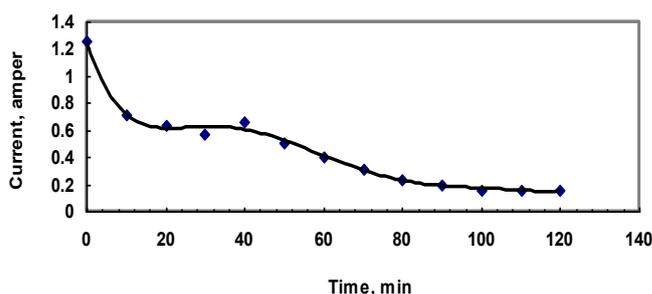


Fig. (2): Current versus time under the following conditions:
 -Applied potential: 2.8 V, pH: 13, NaCl conc.: 90 g/dm³, Initial temp.: 25°C
 -No application of US, PR, or stirring

This layer was ZnO only as indicated by the results of X-Ray Diffraction (XRD). The analysis is shown in Table (1). No trace of zinc hydroxide chloride was observed from the XRD data for the current passivation layer. Usually in the literature, for alkaline solutions, the passivation layer consists predominately of zinc hydroxide chloride or zinc oxide [19,53].

Table (1): XRD analysis of the white material formed in the cell

Observed d-spacing	Reference d-spacing	Phase
2.81906	2.8144	ZnO
2.60635	2.6033	ZnO
2.47825	2.490	ZnO
1.91309	1.9051	ZnO
1.62629	1.6208	ZnO
1.4786	1.4723	ZnO
1.37968	1.3745	ZnO
1.35911	1.3549	ZnO

3.2 Effects of simultaneous application of PR, US and stirring

In Table (2), the effects of simultaneous application of PR, US and stirring on the dissolved zinc and CE are presented. These effects, especially using US, help in minimizing the effect of passivation [57]. In all experiments, the operating parameters were kept the same as indicated in the title of the table.

Table (2): Effects of simultaneous application of PR, US and stirring

Operating conditions:
 - Applied voltage = 2.8 V - NaCl conc. = 90 g/l
 - pH = 13 - Initial temp. = 25 °C

Mode of operation	Dissolved Zn, g	Average current, A	Current efficiency
-No US, PR or stirring	1.0634	0.43625	99.94%
-No US or stirring - use of PR every 2 min.	1.7425	1.349	52.936%
- No stirring or PR - Use of US	3.606	1.5741	93.92%
- No stirring - Use of US & PR every 2 min.	2.4279	1.7017	58.497%
- No PR - Use of US & stirring	3.929	1.616	99.68%

When the polarity was reversed every 2 minutes, without the application of US enhancement or stirring, the mass of dissolved zinc was 1.7425 g and the average current was 1.349 A. Clearly, the application of polarity reversal led to more dissolution of zinc as the passing current was not steadily decreased as in the case of fig. (2). This might be due to the smaller thickness of the solid layer formed at the surface of the anode by the effect of polarity reversal. However, the color of produced solid was slightly darker than the color of the solid produced in the run of fig. (2).

At the end of the experiment, it was observed that, a black layer was formed at the surface of the cathode. The formation of this black layer on zinc electrode was also reported by Cachet et al [39]. When this layer was peeled off the cathode and left for some time in the atmospheric air, it was turned to light grey color. The result of XRD analysis of this grey powder shown in Table (3) indicated the presence of free zinc metal and zinc oxide. The existence of zinc with zinc oxide was also observed by other investigator [58]. The CE was unexpectedly low and had a value of 52.94%. This low CE may be attributed to the fact that there was a time lag between the commencement of polarity reversal and the recommencement of the dissolution process. During this time lag no gases were observed to evolve at the cathode.

Table (3): XRD analysis of the grey layer formed at the cathode

Observed d-spacing	Reference d-spacing	Phase
2.81443	2.8144	ZnO
2.60151	2.6033	ZnO
2.47558	2.473	Zn
2.09004	2.091	Zn
1.91084	1.9111	ZnO
1.62588	1.6249	ZnO
1.47885	1.4772	ZnO
1.3786	1.3784	ZnO
1.36096	1.342	Zn

When US enhancement was only applied at the same operating conditions, the average current was 1.574 A, the dissolved Zn was 3.606 g, and the CE was 93.92%. It was observed that at the end of the experiment the temperature of the bath increased to about 55 °C in spite the initial temperature was about 25 °C. Hence, the application of ultrasonic enhancement led to higher zinc dissolution accompanied with higher CE. Both US and PR every 2 minutes were applied while other operating parameters were kept constant. The average current was estimated as 1.702 A, dissolved Zn was 2.4279 g, and the CE was 58.497%. This indicates that the application of PR had a drawback effect on the dissolution process of Zn. This drawback effect may be a result of the time elapsed before the reaction got started after the reversing of polarity. It may be deduced that PR is not advantageous when used alone or when combined with US.

Finally, during the application of the combination of US and stirring, the average current was 1.616 A, dissolved Zn was 3.929 g, and the CE was 99.68. However the powder was not as white as, for instance, the powder produced in the experiment of Fig. (2). The rise in the temperature of the bath during the run of the experiment was also observed.

3.3 Effect of Current Density (CD)

Fig. (3) indicates that the mass of the dissolved zinc is increasing with the increase in CD. It was not possible to increase CD beyond the largest value shown in the graph due to a limitation imposed by the power supply. Fig. (4), however, indicates that no significant increase of CE was observed with the increase in CD under the conditions of Fig. (3). The CE is sufficiently high enough to confirm that most of the electricity was used to dissolve zinc.

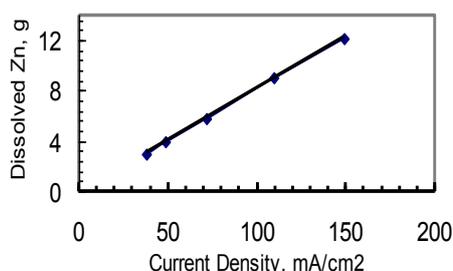


Fig. (3): Effect of current density on anodic dissolution of zinc under below conditions:
 - Application of US and stirring - Cathode is made of stainless steel
 - NaCl concentration = 90 g/l - Starting temperature = 25 °C - pH = 13

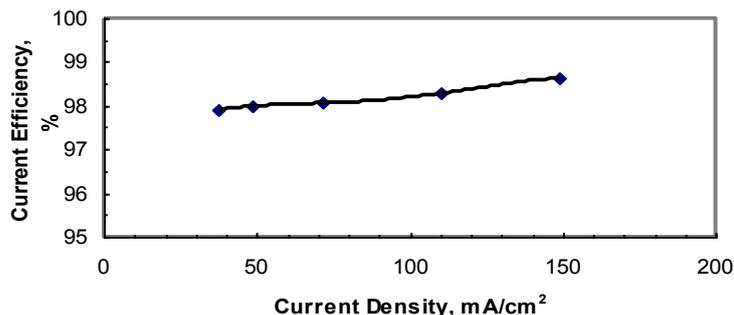


Fig. (4): Effect of current density on current efficiency

3.4 Effect of supporting electrolyte concentration

Fig. (5) and Fig. (6) illustrate the effect of concentration of supporting electrolyte on the amount of dissolved Zinc and CE, respectively. The experiments were performed under the following conditions: application of US, stainless steel cathode, pH = 13, starting applied potential = 5.5 V, and starting temperature = 25 °C. The amount of dissolved zinc, as shown in fig. (5), increased as the concentration of NaCl was increased from 30-150 g/dm³. This may be a result of the increased concentration of chloride ions, which are aggressive ions to the passive layer formed at the anode surface, with respect to hydroxide ions that are inhibitive ions, i.e. lead to the formation of the passive layer [4,17]. Therefore at high concentration of chloride ions, the breakdown of the passive layer is easier and new surface of zinc anode was subjected to dissolution [17]. On the other hand the current efficiency was essentially the same and had a value around 98%, as shown in fig. (6). During these experiments it was observed that the passing current, i.e. the current density, increased with increasing concentration of the supporting electrolyte at the same applied potential. This may be a result of the increase in conductivity of the electrolyte with increasing the concentration of NaCl.

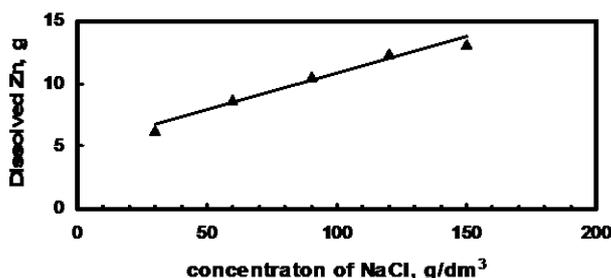


Fig. (5): Effect of the concentration of supporting electrolyte on anodic dissolution of zinc.

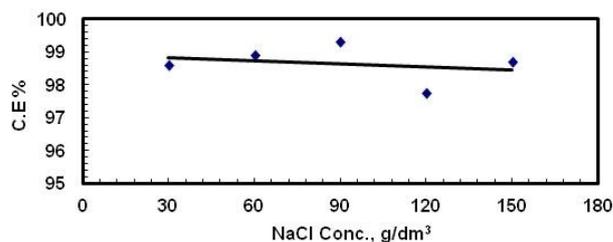


Fig. (6): Effect of NaCl concentration on CE.

3.5 Effect of pH of the bath

In order to investigate the effect of pH of the bath on the dissolution of zinc, four experiments were performed. Values of the pH of the bath were 7, 9, 11, and 13, while the following operating conditions were kept constant: concentration of NaCl = 120 g/dm³, applied potential = 5.5V, application of US, stainless steel cathode, and time of electrolysis = 120 min.

Fig. (7) indicates that the dissolved zinc slightly decreased with increased pH of the bath. This decrease may be a result of the amphoteric nature of the produced ZnO. Fig. (8) shows that the current efficiency was almost constant with increased pH. These results indicates that pH of the bath did not play a significant effect on the amount of the dissolved zinc or current efficiency at the range of pH covered. This may be approved by consulting Pourbaix diagram of zinc; as at the pH value of 7, zinc oxide starts to form at a

potential of 0.8 V. The value of the potential, at which zinc oxide starts to form, decreases with increasing the pH. Since the following work was carried out at higher values of potential than 0.8 V, the effect of pH of the bath is minor.

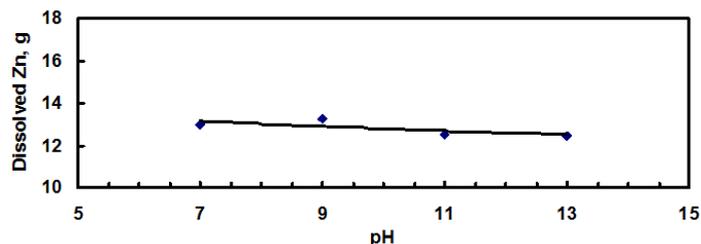


Fig. (7): Effect of pH of the bath on the amount of zinc dissolved.

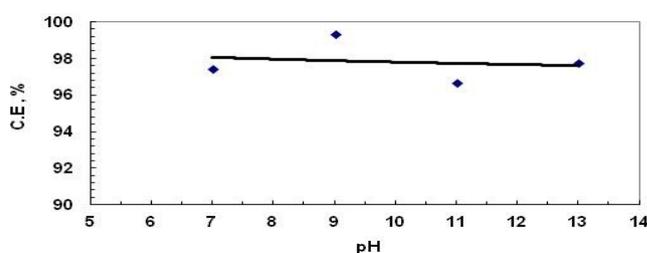


Fig. (8): Effect of pH of the bath on CE.

3.6 Effect of Solution Temperature

During the performance of experiments with the application of US waves, it was observed that at the end of an experiment the temperature of the electrolyte raised to a higher value than the starting temperature, which was about 25 °C. The final temperature was about 55 °C when the applied voltage was 5.5 V. Therefore, it was necessary to investigate the effect of the temperature of the bath on anodic dissolution of zinc without the application of US waves.

The electrolytic cell was placed in a water bath and the temperature was adjusted either by heating or cooling. Cooling was done by adding ice and the temperature was closely monitored. Fig. (9) and Fig. (10) indicate that both the amount of dissolved zinc and CE increased with increasing temperature with the larger effect is on the mass of dissolved zinc. Increasing the temperature from 10 °C to 40 °C led to a significant increase in the mass of the dissolved zinc; but the increase of temperature from 40 to 60 °C, however, did not have a significant effect. This could be attributed to gas passivation.

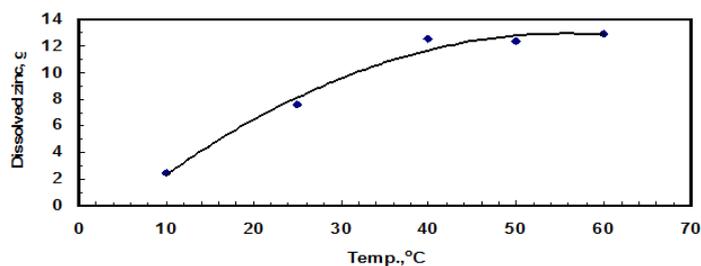


Fig. (9): Effect of temperature on dissolved zinc.

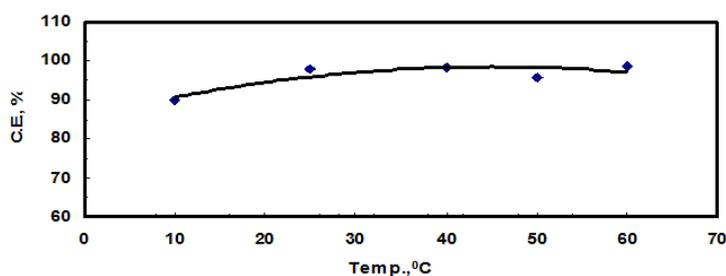


Fig. (10): Effect of temperature on CE.

IV. Conclusions

1. The application of polarity reversal led to more dissolution of zinc but the current efficiency was low.
2. The application of ultrasonic enhancement led to higher zinc dissolution accompanied with higher CE
3. The combination of ultrasonic enhancement and stirring led to more dissolution of zinc.
4. Increasing the current density increased the dissolution of zinc and the current efficiency was almost constant.
5. Increasing concentration of NaCl increased the dissolution of zinc, the current efficiency was almost constant.
6. pH of the bath did not play a significant effect on the amount of the dissolved zinc or current efficiency
7. Increasing the temperature from 10 °C to 40 °C led to a significant increase in the mass of the dissolved zinc and CE; however, the increase of temperature from 40 °C to 60 °C did not have a significant effect.

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